Name: ­­\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Per: \_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_

Classifying Chemical Reactions

More Practice! - Answers

**Part I:**

Classify each of the following reactions as Synthesis (S), Decomposition (D), Single Replacement (SR), Double Replacement (DR), or Combustion (C).

1. \_\_\_DR\_\_\_ BaCl2 + Na2 SO4 → 2 NaCl + BaSO4
2. \_\_\_C\_\_\_\_ C3H8 + 5O2 → 3 CO2 + 4 H2O
3. \_\_\_DR\_\_\_ 3 NH3 + 3I2 → 2 NI3 + 3 H2
4. \_\_\_\_D\_\_\_ 2 H2O2 → 2 H2O + O2
5. \_\_\_SR\_\_\_ Cl2 + 2 NaBr → 2 NaCl + Br2

**Part II:**

Predict the products of the following reactions. Remember to make sure you balance each equation. Reminders: oxidation states, diatomic vs. monatomic molecules, and solubility rules.

1. S8 + 24 F2 → 8 SF6
2. HNO3 + KOH → KNO3 + H2O
3. 2 Al2O3 → 4 Al + 3 O2
4. CuO +H2 → Cu + H2O
5. PbSO4 + 2 AgNO3 → Ag2SO4 + Pb(NO3)2